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Chem 12 Notes On Acids

Chemistry 12 Notes on Unit 4 –

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-Unit 4 -Notes Page 2 - Any acid (weak or strong) could have high or low concentration. Weak & Strong α refers to % ionization. Concentration α the moles of acid dissolved per litre. Eg.)
10.0 M HCl α conc. and strong $[H_3O^+]$
= 10.0 M 0.001 M HCl α dilute and strong $[H$

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Chem 12- Notes on Acids & Bases -
SSS Chemistry

Acids from HClO_4 to H_2SO_4 are 100% ionized in water . only solvent used in Chem 12 (and most Chemistry) so even though HClO_4 is above HCl on

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the chart, it is no more acidic in a water solution. H_3O^+ is the strongest acid that can exist in an undissociated form in water solution. all stronger acids ionize to form H_3O^+

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SSS Chemistry

In this chapter, we will focus on acids and bases and their chemistry. 12.1: Introduction Certain household chemicals, such as some brands of cleanser, can be very concentrated bases, which makes them among the most potentially hazardous

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Chemistry substances found around the home; if spilled on the skin, the strong caustic compound can immediately remove H^+ ions from the flesh, resulting in chemical burns.

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-Unit 4 -Notes Page 2 - Any acid (weak or strong) could have high or low concentration. Weak & Strong refers to % ionization. Concentration refers to the

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moles of acid dissolved per litre. Eg.)
10.0 M HCl à conc. and strong [H³O⁺]
=

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-Unit 4 -Notes Page 2 - Any acid (weak or strong) could have high or low concentration. Weak & Strong refers to % ionization. Concentration refers to the moles of acid dissolved per litre. Eg.)
10.0 M HCl is conc. and strong [H³O⁺]
= 10.0 M 0.001 M HCl is dilute and

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strong [H Chem 12- Notes on Acids & Bases - SSS Chemistry 12.4: Acid-Base Titrations A titration is the quantitative reaction of an acid and a base.

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12 Notes on Unit 4 – Acids, Bases and

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Page 2 - Any acid (weak or strong)

could have high or low concentration.

Weak & Strong à refers to %

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Chemistry. Concentration à the moles of acid dissolved per litre. Eg.) 10.0 M HCl à conc.

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EQUILIBRIA (Unit 4 p 1-12) Strong

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Chemistry 12

Acids and Bases - Section 14 of
General Chemistry Notes is 23 pages
in length (page 14-1 through page

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14-23) and covers ALL you'll need to know on the following lecture/textbook topics: SECTION 14 - Acids and Bases 14-1 -- Acid-Base Chemistry · Arrhenius Acids and Bases · Bronsted-Lowry Acids and Bases · Proton (H^+) Donors and Proton (H^+) Acceptors

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Chemistry Notes | Chemistry Pdf --
Acids, Bases, and the ...

1. The properties of acids and bases
 2. pH and pOH calculations
 3. Definitions of acids and bases
 4. Neutralization reactions.
- NOTES:

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Properties of acids : • Form hydrogen ions (H^+ or protons) in solution • Form hydronium (a water molecule bonded to a hydrogen ion shown as H_3O^+ or $H^+ \cdot H_2O$).

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AND 21 Acids and Bases ...

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Bases, and Salts Unit 5

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Theory Of Acids and Bases • Acid:
any substance which releases $H^+ (aq)$
in water. • Base: any substance
which releases $OH^-(aq)$ in water. •
Salt: is the neutralization product
which results when an acid and a
base react: $HCl (aq) + NaOH (aq) \rightarrow NaCl$
 $(aq) + H_2O (l) \dots$

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Chemistry 12

Acids are classified as Organic Acids and Mineral Acids. Acids which are derived from plants and animals, they are known as Organic Acids. For Example, Citric Acid from fruit.

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Mineral acids are inorganic acids such as Sulphuric Acid. They are dangerous to be used, so need more precautions. Acids are also classified as Strong Acids or Weak Acids. Strong acid is an acid, that completely dissociates into ions in aqueous solutions.

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Revision Notes for Science Chapter 2 -
Acids, bases and ...

Acid-Base Titrations. Indicators.

Titration Curves & Solving AB

Titration Problems (Hebden

summary) Acid Rain. Notes from

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**Titration Suggested Questions to
prep for written component of
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69.e), 99, 101, 104, 113, 115, 118, 120,
125.e), 126, 130, 139, 140**

Unit 4 Acids, Bases, and Salts | Grade

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11 & 12 Notes

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CONCISE A* A Level Chemistry

Edexcel Chapter 12 Notes ...

Phosphoric acid (H_3PO_4) etc.

Chemical Properties of Acid: (i)

Reaction of acids with metal: Acids give hydrogen gas along with respective salt when they react with a metal.

$$\text{Metal} + \text{Acid} \rightarrow \text{Salt} + \text{H}_2$$

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Chemistry Examples: Hydrogen gas and zinc chloride are formed when hydrochloric acid reacts with zinc metal.

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Class Notes: I) Weak Acid Equilibrium
and K_a ; II) Weak Base Equilibrium and
 K_b ; III) Writing Formula (Molecular),

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Complete Ionic and Net Ionic
Equations for Acid/Base ...

Chemistry 12 - Miss Zukowski's Class
Notes Chapter - 2 Acids, Bases and
Salts ACIDS: • These are the
substances which have sour in taste.

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- They turn blue litmus solution to red.
 - They give H⁺ ions in aqueous solution
- Examples of Acids : - HCl - Hydrochloric Acid
- BASES**
- These substances are bitter in taste.
 - They turn red litmus solution to blue.

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Chapter -2, Acids Bases and Salts Notes

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